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**Pearson Edexcel  
International GCSE**

Centre Number

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Candidate Number

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# Chemistry

**Unit: 4CH0**

**Paper: 2C**

Wednesday 17 January 2018 – Afternoon

**Time: 1 hour**

Paper Reference

**4CH0/2C**

**You must have:**

Calculator

Total Marks

## Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided – *there may be more space than you need.*
- Show all the steps in any calculations and state the units.
- Some questions must be answered with a cross in a box ☒. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☒.

## Information

- The total mark for this paper is 60.
- The marks for **each** question are shown in brackets – *use this as a guide as to how much time to spend on each question.*

## Advice

- Read each question carefully before you start to answer it.
- Write your answers neatly and in good English.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ►

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THE PERIODIC TABLE

1 2 3 4 5 6 7 0

Group

Period

4 He Helium 2
------------------------

1 H Hydrogen 1
-------------------------

7 Li Lithium 3	9 Be Beryllium 4	11 Na Sodium 11	12 Mg Magnesium 12	13 Al Aluminium 13	14 Si Silicon 14	15 P Phosphorus 15	16 S Sulfur 16	17 Cl Chlorine 17	18 Ar Argon 18								
19 K Potassium 19	20 Ca Calcium 20	21 Sc Scandium 21	22 Ti Titanium 22	23 V Vanadium 23	24 Cr Chromium 24	25 Mn Manganese 25	26 Fe Iron 26	27 Co Cobalt 27	28 Ni Nickel 28	29 Cu Copper 29	30 Zn Zinc 30	31 Ga Gallium 31	32 Ge Germanium 32	33 As Arsenic 33	34 Se Selenium 34	35 Br Bromine 35	36 Kr Krypton 36
37 Rb Rubidium 37	38 Sr Strontium 38	39 Y Yttrium 39	40 Zr Zirconium 40	41 Nb Niobium 41	42 Mo Molybdenum 42	43 Tc Technetium 43	44 Ru Ruthenium 44	45 Rh Rhodium 45	46 Pd Palladium 46	47 Ag Silver 47	48 Cd Cadmium 48	49 In Indium 49	50 Sn Tin 50	51 Sb Antimony 51	52 Te Tellurium 52	53 I Iodine 53	54 Xe Xenon 54
55 Cs Caesium 55	56 Ba Barium 56	57 La Lanthanum 57	72 Hf Hafnium 72	73 Ta Tantalum 73	74 W Tungsten 74	75 Re Rhenium 75	76 Os Osmium 76	77 Ir Iridium 77	78 Pt Platinum 78	79 Au Gold 79	80 Hg Mercury 80	81 Tl Thallium 81	82 Pb Lead 82	83 Bi Bismuth 83	84 Po Polonium 84	85 At Astatine 85	86 Rn Radon 86
87 Fr Francium 87	88 Ra Radium 88	89 Ac Actinium 89	101 Fe Iron 101	102 Co Cobalt 102	103 Rh Rhodium 103	104 Pd Palladium 104	105 Ag Silver 105	106 Ni Nickel 106	107 Cu Copper 107	108 Zn Zinc 108	109 Ga Gallium 109	110 Ge Germanium 110	111 As Arsenic 111	112 Se Selenium 112	113 Br Bromine 113	114 Kr Krypton 114	

Key

Relative atomic mass
Symbol
Name
Atomic number

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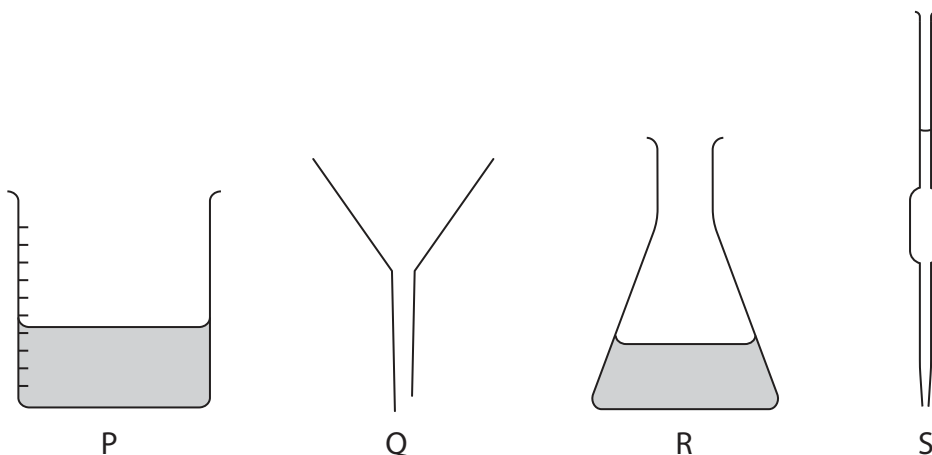
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Answer ALL questions.

1 These pieces of apparatus are used in chemistry experiments.



(a) Name these pieces of apparatus.

(4)

P .....

Q .....

R .....

S .....

(b) Apparatus P contains dilute hydrochloric acid.

Litmus indicator is added to this acid.

What is the final colour of the litmus?

(1)

- A blue       B green       C orange       D red

(c) Apparatus R contains potassium hydroxide solution.

Litmus indicator is added to this alkaline solution.

What is the final colour of the litmus?

(1)

- A blue       B green       C orange       D red

(Total for Question 1 = 6 marks)

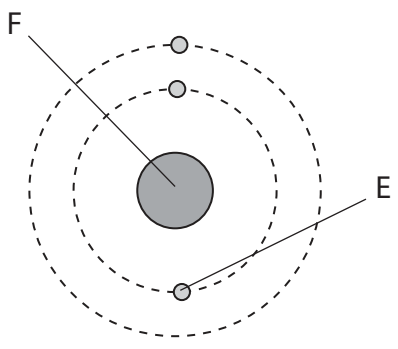
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2 The diagram shows an atom of lithium with atomic number 3 and mass number of 6.



(a) Name the particle labelled E. (1)

.....

(b) Name the part of the atom labelled F. (1)

.....

(c) Name the two types of particle found in part F. (2)

1 .....

2 .....



(d) Another type of lithium atom has atomic number 3 and mass number 7.

- (i) State the name given to atoms with the same atomic number but different mass numbers.

(1)

- (ii) Draw a diagram to show the arrangement of electrons in an atom of lithium with atomic number 3 and mass number 7.

(1)

(e) A sample of lithium contains 92.5% of atoms with mass number 7 and 7.5% of atoms with mass number 6.

Calculate the relative atomic mass of lithium.

(2)

relative atomic mass = .....

**(Total for Question 2 = 8 marks)**

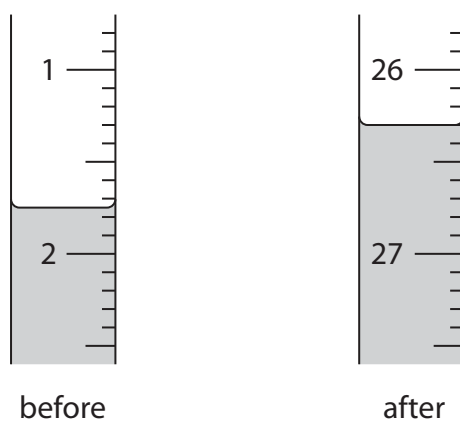


3 A student is provided with a solution of dilute sulfuric acid and a solution of sodium hydroxide. The student does a titration using  $25.0 \text{ cm}^3$  of the sodium hydroxide solution. She adds the acid from a burette.

(a) Which type of reaction occurs between dilute sulfuric acid and sodium hydroxide? (1)

- A displacement
- B neutralisation
- C precipitation
- D redox

(b) The diagram shows the student's burette readings for the titration.



(i) Use the readings to complete the table, giving all values to the nearest  $0.05 \text{ cm}^3$ . (3)

burette reading after adding acid	
burette reading before adding acid	
volume in $\text{cm}^3$ of acid added	



(ii) Explain why the student needs to repeat the titration in order to obtain a reliable value for the volume of acid required to react exactly with 25.0 cm<sup>3</sup> of sodium hydroxide solution.

(2)

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**(Total for Question 3 = 6 marks)**

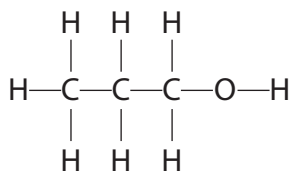
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- 4 (a) This is the displayed formula of an organic compound, X.



- (i) Give the molecular formula of compound X. (1)

- (ii) A student describes compound X as a saturated hydrocarbon.  
Explain whether the student is correct. (3)

- (b) Compound X and ethanol are members of the homologous series of alcohols.

One property of members of a homologous series is that they have similar chemical reactions.

- Give one other property of members of a homologous series. (1)





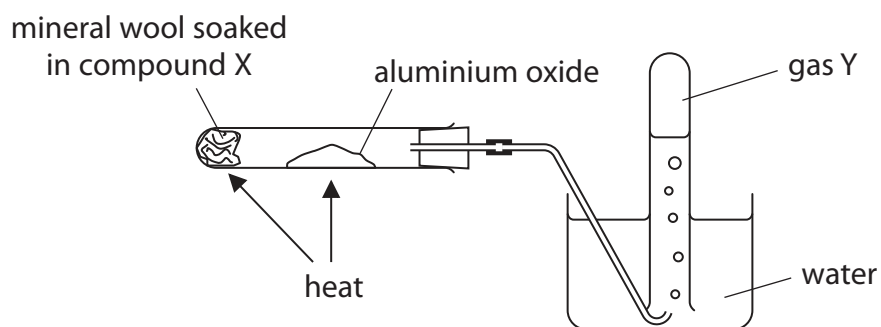
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(c) This apparatus is used for a dehydration reaction using compound X.

This reaction is similar to the dehydration reaction of ethanol.



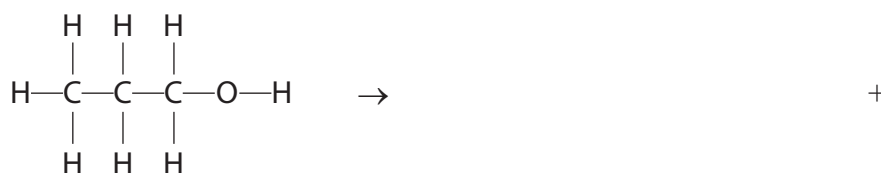
The reaction produces gas Y and one other product.

(i) State the purpose of the aluminium oxide. (1)

(ii) State a property of gas Y that allows it to be collected over water. (1)

(iii) Give a reason why the first sample of gas Y collected is not pure. (1)

(iv) Complete the equation for the dehydration reaction showing the displayed formula of gas Y and the molecular formula of the other product. (2)



(v) Give the name of gas Y. (1)

(Total for Question 4 = 11 marks)



5 Chromium is a shiny metal that has many uses.

Most chromium is extracted from the ore chromite,  $\text{FeCr}_2\text{O}_4$

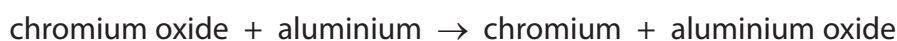
(a) Complete the table by giving the names of the elements in  $\text{FeCr}_2\text{O}_4$

(1)

Chemical symbol	Name of element
Fe	
Cr	
O	

(b) In the extraction process, chromite is converted into chromium(III) oxide,  $\text{Cr}_2\text{O}_3$

Chromium is made by this reaction



(i) Write a chemical equation for this reaction.

(2)

(ii) Explain what the reaction shows about the reactivity of chromium compared to the reactivity of aluminium.

(2)

(iii) Explain why the reaction between chromium oxide and aluminium is described as a redox reaction.

(2)



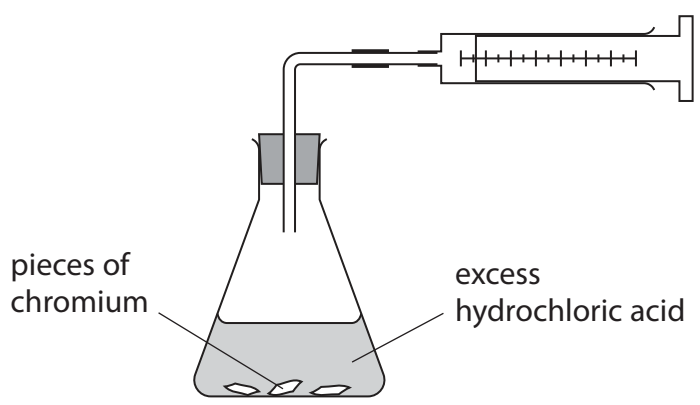
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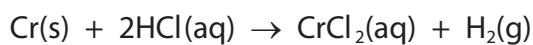
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(c) Chromium metal reacts with dilute hydrochloric acid to form hydrogen gas.

This apparatus is used to investigate the reaction.



The equation for the reaction is



A student adds 0.13 g of a sample of chromium metal to excess dilute hydrochloric acid.

- (i) Calculate the maximum volume of hydrogen gas that the student could produce in this experiment at room temperature and pressure (rtp).

[molar volume of a gas is 24 dm<sup>3</sup> at rtp]

(3)

maximum volume = ..... dm<sup>3</sup>

- (ii) The student does the experiment at rtp and finds that the volume collected is less than the calculated maximum.

Give two possible reasons for this.

(2)

1 .....

2 .....

(Total for Question 5 = 12 marks)



6 Lithium fluoride, LiF, and magnesium oxide, MgO, are ionic compounds.

(a) (i) Calculate the relative formula mass ( $M_r$ ) of MgO.

(1)

$M_r =$  .....

(ii) Give the formulae of the two ions in LiF.

(1)

..... and .....

(b) Explain why

- ionic compounds have high melting points
- the melting point of magnesium oxide is much higher than the melting point of lithium fluoride

(4)

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(c) Explain why ionic compounds do not conduct electricity when solid, but do conduct electricity when molten or in aqueous solution.

(2)

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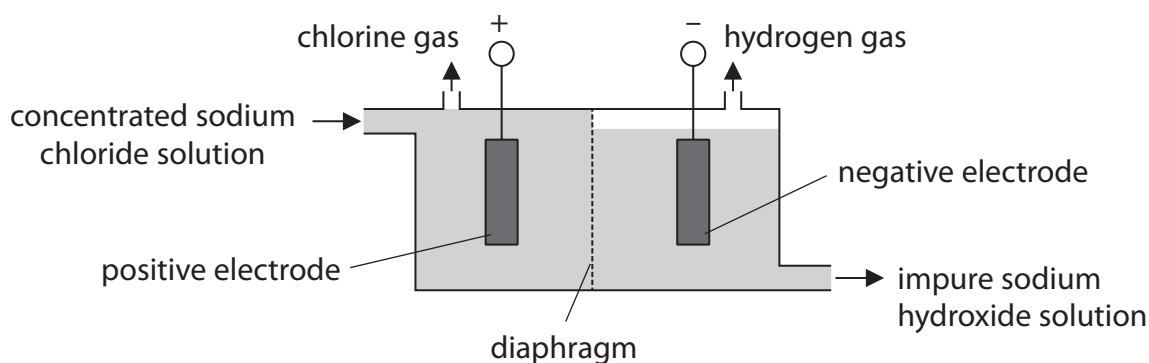
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**(Total for Question 6 = 8 marks)**



7 The diagram shows the electrolysis of concentrated sodium chloride solution in a diaphragm cell.



(a) (i) The ionic half-equation for the reaction at the positive electrode is



Use this equation to explain why oxidation occurs at the positive electrode.

(2)

.....

.....

.....

.....

(ii) At the negative electrode, water molecules gain electrons to form hydroxide ions and hydrogen gas.

Complete the ionic half-equation for this reaction.

(2)



(b) Chlorine reacts with sodium hydroxide to produce a mixture of water, sodium chloride and sodium chlorate(I), NaOCl.

Write a chemical equation for this reaction.

(1)

.....



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(c) Chlorine is used in the manufacture of the addition polymer poly(chloroethene).

(i) Explain how an addition polymer is formed from its monomers. (2)

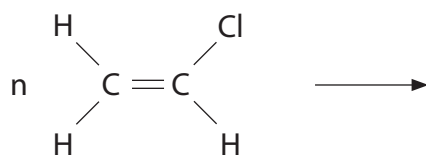
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(ii) Complete this equation by drawing the displayed formula of poly(chloroethene). (2)



(Total for Question 7 = 9 marks)

TOTAL FOR PAPER = 60 MARKS



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